Cambridge
O Level

Cambridge International Examinations
Cambridge Ordinary Level

CHEMISTRY **5070/11**

Paper 1 Multiple Choice **October/November 2014**

1 hour

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB recommended)

1792508481

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.
Do not use staples, paper clips, glue or correction fluid.
Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.
DO **NOT** WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.
Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.
Any rough working should be done in this booklet.
A copy of the Periodic Table is printed on page 16.
Electronic calculators may be used.

This document consists of **14** printed pages and **2** blank pages.

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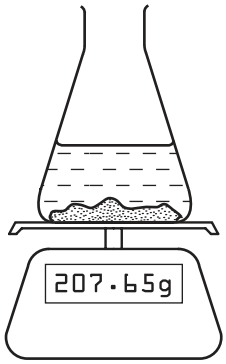
1 Calcium carbonate reacts with hydrochloric acid, producing carbon dioxide gas.

$$CaCO_3(s) + 2HCl(aq) \rightarrow CaCl_2(aq) + H_2O(l) + CO_2(g)$$

The rate of this reaction can be measured using the apparatus shown.



Which additional piece of apparatus is also required?

A a burette

B a clock

C a gas syringe

D a thermometer

2 Which compound when in aqueous solution will produce a red / brown precipitate on the addition of an aqueous solution of Fe^{3+} ions?

A hydrogen chloride

B sodium chloride

C sodium hydroxide

D sulfur trioxide

3 What is the correct sequence for obtaining pure salt from a mixture of sand and salt?

A add water, evaporate

B add water, filter

C add water, filter, evaporate

D filter, add water, evaporate

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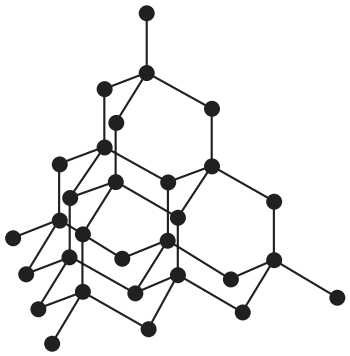
4 The diagram shows the structure of which element in Period 3?



A aluminium

B magnesium

C silicon

D sodium

5 The table contains information on the structure of four particles.

particle	proton number	number of protons	number of neutrons	number of electrons
Mg	12	12	W	12
Mg^{2+}	12	12	12	X
F	Y	9	10	9
F^-	9	9	10	Z

What are the values of W, X, Y and Z in the table above?

	W	X	Y	Z
A	10	12	9	10
B	12	10	9	10
C	12	10	10	9
D	12	12	10	9

6 Which statement describes ionic bonding?

A a lattice of ions in a sea of electrons

B electrostatic attraction between oppositely charged ions

C the sharing of electrons between atoms to gain a noble gas configuration

D the transfer of electrons from atoms of a non-metal to the atoms of a metal

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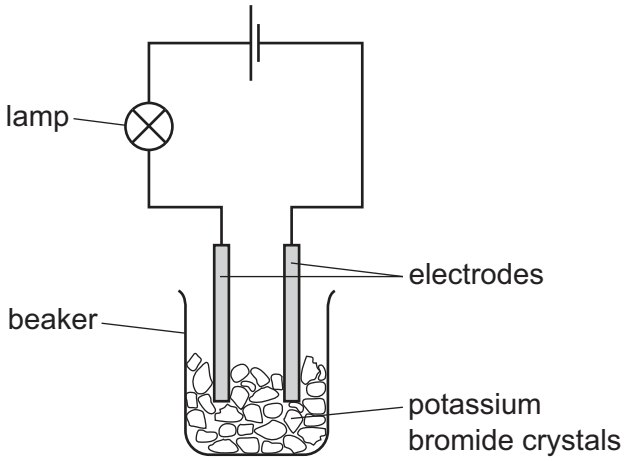
7 The experiment shown is used to test potassium bromide crystals.



The lamp does not light.

Distilled water is then added to the beaker and the lamp lights.

Which statement explains these results?

A Electrons are free to move in the solution when potassium bromide dissolves.

B Metal ions are free to move when potassium bromide melts.

C Metal ions are free to move when potassium reacts with water.

D Oppositely charged ions are free to move in the solution when potassium bromide dissolves.

8 Why does ammonia gas diffuse faster than hydrogen chloride gas?

A Ammonia has a higher boiling point than hydrogen chloride.

B Ammonia is a base, hydrogen chloride is an acid.

C The ammonia molecule contains more atoms than a hydrogen chloride molecule.

D The relative molecular mass of ammonia is smaller than that of hydrogen chloride.

9 Which molecule has only four electrons involved in covalent bonds?

A H_2S **B** CO_2 **C** Cl_2 **D** N_2

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10 A volume of ethane, C_2H_6, at r.t.p. has a mass of 20 g.

What is the mass of an equal volume of propene, C_3H_6, at r.t.p.?

A 20 g **B** 21 g **C** 28 g **D** 42 g

11 Which element requires the largest number of electrons for one mole of the metal to be formed from its aqueous ions during electrolysis?

A aluminium

B calcium

C copper

D sodium

12 Which changes are observed during the electrolysis of aqueous copper(II) sulfate using copper electrodes?

1 A pink solid is deposited on the negative electrode.

2 Bubbles form on the positive electrode.

3 The colour of the solution does not change.

A 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3

13 Analysis of a sample of an oxide of nitrogen gave the following data.

- percentage by mass of nitrogen 47%
- percentage by mass of oxygen 53%

What is the empirical formula of this oxide?
[A_r: N, 14; O, 16]

A NO **B** NO_2 **C** N_2O **D** N_2O_3

14 Petroleum is a mixture of hydrocarbons which can be separated into fractions by fractional distillation.

Which row shows the fractions in order of decreasing boiling point?

	highest b.p. ⟶			lowest b.p.
A	diesel	paraffin	naphtha	petrol
B	paraffin	naphtha	petrol	diesel
C	naphtha	petrol	diesel	paraffin
D	petrol	naphtha	paraffin	diesel

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15 Which is **not** true about the process of photosynthesis?

A Carbon dioxide and water react in a 1 : 1 molar ratio.

B Glucose is produced and can be used as a source of energy.

C Oxygen is produced.

D The reaction is exothermic.

16 The equation shows the reaction for the manufacture of ammonia.

$$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$$

Which change will decrease the activation energy of the reaction?

A addition of a catalyst

B decrease in temperature

C increase in concentration

D increase in pressure

17 Which ionic equation represents a redox reaction?

A $Ag^+ + Cl^- \rightarrow AgCl$

B $Ba^{2+} + SO_4^{2-} \rightarrow BaSO_4$

C $H^+ + OH^- \rightarrow H_2O$

D $Zn + Cu^{2+} \rightarrow Zn^{2+} + Cu$

18 The equation shows the reaction for the formation of sulfur trioxide using a catalyst.

$$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g) \qquad \Delta H = -197\,\text{kJ/mol}$$

Which change in reaction conditions would produce more sulfur trioxide?

A adding more catalyst

B decreasing the pressure

C increasing the temperature

D removing some sulfur trioxide

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19 To which substance is dilute sulfuric acid added to prepare lead(II) sulfate?

A aqueous lead(II) nitrate

B lead foil

C powdered lead(II) carbonate

D powdered lead(II) oxide

20 Which metal can react with water at r.t.p.?

A calcium

B copper

C lead

D zinc

21 Which statement about amphoteric oxides is **not** correct?

A They dissolve in water.

B They are formed only by metals.

C They react with aqueous sodium hydroxide to give salts.

D They react with aqueous acids to give salts.

22 Which statement explains why the chemical properties of sodium and potassium are similar?

A They are in the same group of the Periodic Table.

B They are in the same period of the Periodic Table.

C They are soft and can be cut with a knife.

D They have similar melting points.

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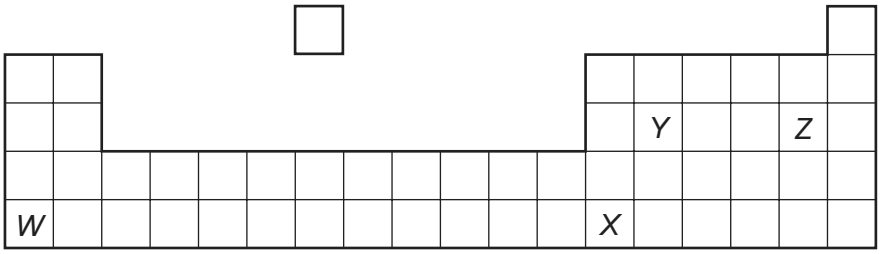
23 The diagram shows an outline of part of the Periodic Table.



Which statement is **not** correct?

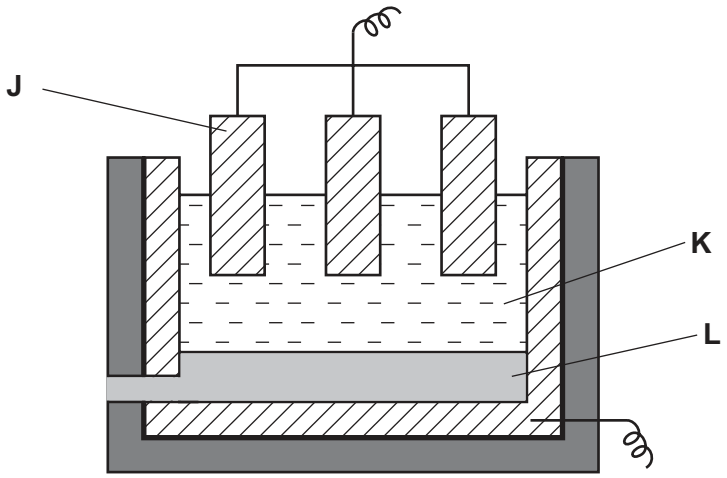
A The melting point of *W* is lower than that of *Z*.

B *W* and *Z* could react together and form a compound, *WZ*.

C *X* could form an oxide, X_2O_3.

D *Y* could form an oxide, YO_2.

24 The diagram shows apparatus that can be used to extract aluminium.



What are **J**, **K** and **L**?

	J	K	L
A	negative electrode	aluminium oxide + cryolite	aluminium
B	negative electrode	cryolite	aluminium oxide
C	positive electrode	aluminium oxide	cryolite
D	positive electrode	aluminium oxide + cryolite	aluminium

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25 Sulfur is burnt in air.

Which statement about this reaction is correct?

A The gas formed turns aqueous potassium dichromate(VI) from green to orange.

B The product is used as a food preservative.

C The reaction is endothermic.

D The reaction is reversible.

26 A gas **G**

1 has no smell,

2 is not poisonous,

3 reacts with hydrogen at high temperature and pressure.

What is gas **G**?

A carbon monoxide

B helium

C nitrogen

D chlorine

27 Which method of water purification can be used to obtain drinkable water from seawater?

A chlorination

B desalination

C filtration

D sedimentation

28 Which atmospheric pollutant is produced by bacterial decay of vegetable matter?

A carbon monoxide

B methane

C ozone

D sulfur dioxide

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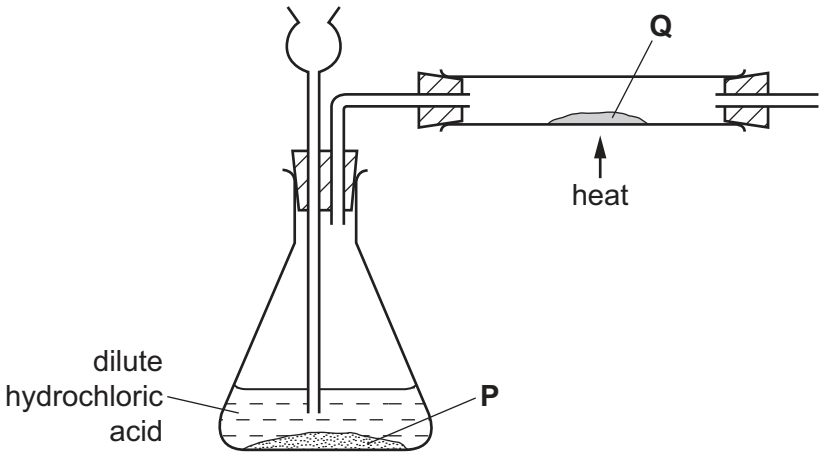
29 Substance **P** reacts with dilute hydrochloric acid to produce a gas.
This gas reduces substance **Q**.



What are substances **P** and **Q**?

	P	Q
A	copper	copper(II) oxide
B	lead	lead(II) oxide
C	magnesium	zinc oxide
D	zinc	copper(II) oxide

30 Which two statements about alloys are correct?

1 Alloys are formed by mixing two metals.

2 Alloys do **not** conduct electricity.

3 Atoms in an alloy must all be the same size.

4 In an alloy there is metallic bonding.

A 1 and 2 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

31 A powdered mixture of metals contains aluminium, calcium, silver and iron. Excess hydrochloric acid is added until no more mixture dissolves.

What is the undissolved residue?

A aluminium

B calcium

C iron

D silver

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32 Iron rusts when exposed to oxygen in the presence of water.

Which method will **not** slow down the rate of rusting of an iron roof?

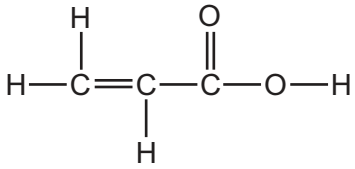
A attaching strips of copper to it

B coating it with plastic

C galvanising it with zinc

D painting it

33 A compound has the following structure.



Which reactions will occur with this compound?

1 Bromine water will decolourise.

2 It will react with an alcohol to form an ester.

3 It will react with sodium metal.

A 1 only **B** 1 and 2 only **C** 1, 2 and 3 **D** 2 and 3 only

34 In the Periodic Table, how many periods are needed to accommodate the elements of atomic numbers 1-18?

A 2 **B** 3 **C** 4 **D** 8

35 A compound **X** has the molecular formula $C_4H_8O_2$. It reacts with calcium carbonate to give carbon dioxide.

What is **X**?

A $HCO_2C_3H_7$

B $CH_3CO_2C_2H_5$

C $C_2H_5CO_2CH_3$

D $C_3H_7CO_2H$

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36 Methane is the first member of the alkane series of hydrocarbons. The second member is ethane.

Which statements about ethane are correct?

1 Ethane has the formula C_2H_4.

2 Ethane has a higher boiling point than that of methane.

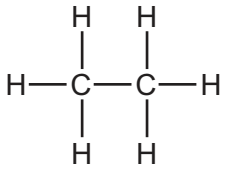
3 Ethane has the same molecular formula as methane.

4 Ethane has chemical properties very similar to those of methane.

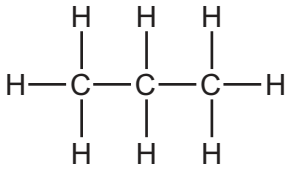
A 1, 2 and 3 **B** 1 and 4 **C** 2 and 4 **D** 3 only

37 Which alkane, when any one hydrogen atom is substituted by a chlorine atom, will **not** produce isomers?

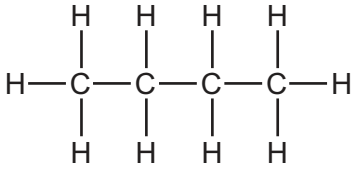
A



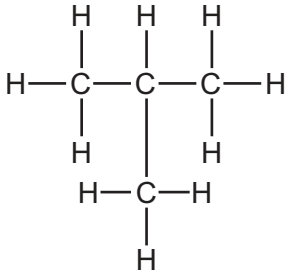
B



C



D



38 When ethanol reacts with ethanoic acid, the ester ethyl ethanoate is formed.

$$C_2H_5OH + CH_3CO_2H \rightarrow CH_3CO_2C_2H_5 + H_2O$$

What is the formula of the ester formed when methanol reacts with butanoic acid, $C_3H_7CO_2H$?

A $C_2H_5CO_2C_2H_5$

B $C_3H_7CO_2C_2H_5$

C $CH_3CO_2C_3H_7$

D $C_3H_7CO_2CH_3$

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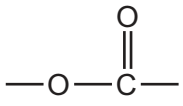
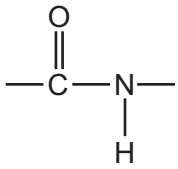
39 The table gives some statements about some macromolecules.

1	fats contain the linkage 	proteins contain the linkage 
2	poly(ethene) is made by addition polymerisation	*Terylene* is made by condensation polymerisation
3	starch can be hydrolysed to produce sugars	proteins can be hydrolysed to produce amino acids
4	*Terylene* is a naturally occurring polymer	nylon is a man-made polymer

Which pairs of statements are correct?

A 1 and 2 only **B** 2 and 3 only **C** 3 and 4 **D** 1, 2 and 3

40 Which of these compounds could react together to form a polymer?

1 $H_2N(CH_2)_6NH_2$

2 $CH_3(CH_2)_4COOH$

3 $HOOC(CH_2)_4COOH$

4 $H_2N(CH_2)_6CH_3$

A 1 and 2 **B** 1 and 3 **C** 2 and 4 **D** 3 and 4

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DATA SHEET
The Periodic Table of the Elements

Group																	
I	II									III	IV	V	VI	VII	0
															4 He Helium 2
					1 H Hydrogen 1										
7 Li Lithium 3	9 Be Beryllium 4									11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10
23 Na Sodium 11	24 Mg Magnesium 12									27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulfur 16	35.5 Cl Chlorine 17	40 Ar Argon 18
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	Tc Technetium 43	101 Ru Ruthenium 44	103 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	128 Te Tellurium 52	127 I Iodine 53	131 Xe Xenon 54
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57 *	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	Po Polonium 84	At Astatine 85	Rn Radon 86
Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89 †															

140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	Pm Promethium 61	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71
232 Th Thorium 90	Pa Protactinium 91	238 U Uranium 92	Np Neptunium 93	Pu Plutonium 94	Am Americium 95	Cm Curium 96	Bk Berkelium 97	Cf Californium 98	Es Einsteinium 99	Fm Fermium 100	Md Mendelevium 101	No Nobelium 102	Lr Lawrencium 103

*58-71 Lanthanoid series
†90-103 Actinoid series

Key: a = relative atomic mass; X = atomic symbol; b = proton (atomic) number

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).

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